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Chapter 22 Hydrocarbon Compounds 237
SECTION 22.1 HYDROCARBONS (pages 693–701) This section describes the bonding in hydrocarbons and distinguishes straight-chain from branched-chain alkanes. It also provides rules for naming branch-chained alkanes. Organic Chemistry and Hydrocarbons (pages 693–694) 1.

SECTION 22.1

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Study Guide Key Concepts 22.1 Hydrocarbons • Because carbon has four valence electrons, car-bon atoms always form four covalent bonds. • The carbon

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atoms in an alkane can be arranged in a straight chain or in a chain that has branches. • Molecules of hydrocarbons, such as alkanes, are nonpolar molecules. 22.2 Unsaturated Hydrocarbons

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The hydrogenation reaction is possible only for compounds with double or triple bonds, and such compounds are said to be unsaturated hydrocarbons. The addition of the hydrogen to the carbon

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atoms that were double- or triple-bonded converts the unsaturated compound to a saturated hydrocarbon with only single bonds.

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A(n) (9) is a compound in which the acidic hydrogen of an organic acid is replaced by a carbon atom or hydrocarbon chain. A(n) (10) is a compound in which the -OH group of an organic acid is replaced by a nitrogen atom bonded to other atoms. The functional group of such compounds is found in (11). CHAPTER 23 STUDY GUIDE FOR CONTENT MASTERY

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Hydrocarbons are organic molecules that only contain two elements: carbon and hydrogen.

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